

**Physical Science**

DSHS

Name \_\_\_\_\_

Period \_\_\_\_\_

1. Ionic bonds form between \_\_\_\_\_ and \_\_\_\_\_
2. In ionic bonding atoms \_\_\_\_\_ and \_\_\_\_\_ electrons
3. Covalent bond form between \_\_\_\_\_ and \_\_\_\_\_
4. In covalent bonds atoms \_\_\_\_\_ electrons
5. Fill in the table below. Be sure to pay attention to whether it is an ionic or covalent bond!!

	compound formula	type of bond (ionic or covalent)	Name of compound		Name of compound	type of bond (ionic or covalent)	compound formula
1	PI <sub>3</sub>			16	Carbon disulfide		
2	H <sub>2</sub> O <sub>2</sub>			17	Sodium oxide		
3	CH <sub>4</sub>			18	Nickel (II) phosphide		
4	NaF			19	Silicon tetraiodide		
5	SiS <sub>2</sub>			20	dihydrogen disulfide		
6	BeI <sub>2</sub>			21	dichloride monofluoride		
7	F <sub>2</sub>			22	aluminum oxide		
8	N <sub>2</sub> O <sub>2</sub>			23	carbon tetrachloride		
9	SrO			24	phosphorous trifluoride		
10	K <sub>2</sub> O			25	calcium fluoride		
11	PF <sub>3</sub>			26	Iron (III) oxide		
12	HF			27	silicon monoxide		
13	SiF <sub>2</sub>			28	sodium nitride		
14	AlF <sub>3</sub>			29	dihydrogen monoxide		
15	CaO			30	silver sulfide		

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Draw the bonding picture for the following compounds.

**If the bond is ionic, be sure to show:**

1. the transfer of electrons
2. the final picture including brackets, the final position of the electrons and the charge on each ion

**If the bond is covalent, be sure to show**

1. the sharing of electrons by drawing circles around each atom, the electrons that were original to that atom and the ones that it is sharing
2. The final picture including dash lines to represent the shared pairs of electrons and electrons that aren't part of shared pairs represented by dots

